

Effect of the surface-stimulated mode on the kinetics of homogeneous crystal nucleation in droplets

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Abstract.

A kinetic theory of homogeneous crystal nucleation in unary droplets is presented taking into account that a crystal nucleus can form not only in the volume-based mode (with all its facets *within* the droplet) but also in the surface-stimulated one (with one of its facets *at* the droplet surface). The previously developed thermodynamics of surface-stimulated crystal nucleation rigorously showed that if at least one of the facets of the crystal is only partially wettable by its melt, then it is thermodynamically more favorable for the nucleus to form with that facet *at* the droplet surface rather than *within* the droplet. So far, however, the kinetic aspects of this phenomenon had not been studied at all. The theory proposed in the present paper advocates that even in the surface-stimulated mode crystal nuclei initially emerge (as sub-critical clusters) homogeneously in the sub-surface layer, not “pseudo-heterogeneously” at the surface. A homogeneously emerged sub-critical crystal can become a surface-stimulated nucleus due to density and structure fluctuations. This effect contributes to the total rate of crystal nucleation (as the volume-based mode does). An explicit expression for the total per-particle rate of crystal nucleation is derived. Numerical evaluations for water droplets suggest that the surface-stimulated mode can significantly enhance the per-particle rate of crystal nucleation in droplets as large as 10 μm in radius. Possible experimental verification of the proposed theory is discussed.

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1 Introduction

The composition, size, and phases of aerosols and cloud particles determine their radiative and chemical properties^{1,2} thus determining the extent to and manner in which the Earth climate is affected. On the other hand, the composition, size, and phases of atmospheric particles are determined by the rate at and mode in which these particles form and evolve.^{2–4} Water constitutes an overwhelmingly dominant chemical species participating in these atmospheric processes hence a great importance attributed to studying aqueous aerosols and cloud droplets as well as their phase transformations.

Although many phase transitions in aqueous aerosols and cloud droplets occur via heterogeneous nucleation on preexisting solid particles, in a number of important cases atmospheric droplets appear to freeze homogeneously.^{2,3,5} For example, the conversion of supercooled water droplets into ice at temperatures below about -30°C is known to occur homogeneously, mainly because the concentrations of the observed ice particles in the clouds often exceed the number densities of preexisting particles capable of nucleating ice.^{6,7}

Crystallization process in pure systems had long been assumed to initiate within the volume of the supercooled phase.^{8,9} Under that assumption, the rate of crystallization of a droplet is proportional to its volume.^{2,3,8,9} However, using the classical nucleation theory (CNT) based on the capillarity approximation,¹⁰ we recently developed^{11,12} a thermodynamic theory that prescribes the condition under which the surface of a droplet can stimulate crystal nucleation therein so that the formation of a crystal nucleus with one of its facets at the droplet surface is thermodynamically favored over its formation with all the facets *within* the liquid phase. This condition has the form of an inequality which, when satisfied, predicts that crystal nucleation in the droplet occurs mostly in a “surface-stimulated” mode rather than in a “volume-based” one. For both unary¹¹ and multicomponent¹² droplets the inequality coincides with the con-

dition for the partial wettability of at least one of the facets of a crystal nucleus by its own melt.¹³ This effect was experimentally observed for several systems,^{14,15} including water-ice¹⁶ at temperatures at or below 0°C.

Although kinetic factors may play as important a role as thermodynamic ones in determining the mode of crystal nucleation, the partial wettability of a solid by its melt may help to explain why, in molecular dynamics simulations of various kinds of supercooled liquid clusters^{17,18} crystal nuclei appear preferentially very close to the surface. As a result, since smaller clusters have a higher surface-to-volume ratio, nucleation rates in smaller clusters tend to be higher than in the bulk. Hence it is experimentally easier to observe the crystallization of aerosols, having a large collective surface area, than those having a large volume. The analysis^{19,20} of laboratory data provided by various authors also suggests that the nucleation of both ice in supercooled water droplets and nitric acid hydrates in concentrated aqueous nitric acid droplets may initiate at the droplet surface layer. Recent experiments^{21,22} on the heterogeneous freezing of water droplets in both immersion and contact modes have also provided evidence that the rate of crystal nucleation in the contact mode is much higher because the droplet surface may stimulate heterogeneous crystal nucleation in the way similar to the enhancement of the homogeneous process.

It is well known that under otherwise identical thermodynamic conditions the free energy barrier of heterogeneous nucleation is usually much lower than that of homogeneous nucleation,¹³ so it might seem that the idea of surface-stimulated crystal nucleation does not significantly contribute to clarifying the underlying physics of the crystallization phenomenon. In this regard, it should be noted that the surface-stimulated crystallization is *not* a particular case of heterogeneous nucleation. On the contrary, it is a particular case of homogeneous crystal nucleation hence its apparent thermodynamic similarities with heterogeneous nucleation

should be interpreted with due caution. Our analysis in the present paper will show that the kinetics of this process cannot be treated by using the formalism of heterogeneous nucleation on foreign surfaces.

The paper is structured as follows. In section 2 we briefly outline main results^{11,12} concerning the thermodynamics of surface-stimulated crystal nucleation occurring homogeneously. For the sake of simplicity, in this work we consider only unary systems, i.e., pure water droplets, but the generalization to multicomponent droplets can be carried out in the same manner as the unary theory in ref.11 was extended to multicomponent systems in ref.12. A kinetic theory of such a process is presented in section 3. Numerical predictions and possible experimental verification of the model are discussed in Section 4. The results and conclusions are summarized in section 5.

2 Thermodynamics of surface-stimulated crystal nucleation

To determine the conditions under which the surface of a droplet can thermodynamically stimulate its crystallization, it is necessary to consider the formation of a crystal cluster a) within a liquid droplet and b) with one of the crystal facets at the liquid-vapor interface (Figure 1). The criterion for the surface-stimulated crystallization is obtained by comparing the reversible works of formation of a crystal nucleus (critical cluster) in these two case. This was done in the framework of CNT for both unary¹¹ and multicomponent¹² droplets. The main results for unary droplets are outlined in this section.

Assuming the crystallization process to be isothermal and neglecting the density difference between liquid and solid phases, one can say that the total volume V , temperature T , and number of molecules N in the system will be constant.^{13,23} Then the reversible work of formation of a crystal cluster, W , can be found as a change in the Helmholtz free en-

ergy of the system upon its transition from the initial state (vapor+liquid) into the final one (vapor+liquid+crystal).^{11,12}

Consider a crystal cluster (phase s) formed *within* a liquid droplet (phase l) which is surrounded by the vapor (phase v) (see Figure 1). The single superscripts l , s , and v will denote quantities in the corresponding phases, whereas double superscripts will denote quantities at the corresponding interfaces. The reversible work of formation of a crystal of arbitrary shape with n facets is

$$W = \nu[\mu^s(P^s, T) - \mu^l(P^l, T)] - V^s(P^s - P^l) + \sum_{i=1}^{\lambda} \sigma_i^{ls} A_i^{ls}, \quad (1)$$

where ν and $\mu^s(P^s, T)$ are the number of molecules and chemical potential in the solid particle formed within the liquid, P^s being the pressure within the crystal and V its volume; $\mu^l(P^l, T)$ is the chemical potential in the remaining liquid with pressure P^l , σ_i^{ls} and A_i^{ls} are the surface tension and surface area of facet i ($i = 1, \dots, n$) of the crystal particle (anisotropic interfacial energies are crucial in determining the character of the nucleation process). By definition,¹³ the surface tension of the solid is equal to the surface free energy per unit area if the adsorption at the solid-fluid interfaces is negligible, which usually is a reasonable assumption. The pressure in the droplet is related to the pressure P^β in the surrounding vapor via the Laplace equation $P^l = P^v + 2\sigma^{lv}/R$, with σ^{lv} being the droplet surface tension and R the droplet radius (considered constant during crystallization).

Assuming that in the temperature range between T and the bulk melting temperature T_0 the enthalpy of fusion does not change significantly, one can rewrite equation (1) as

$$W = -\nu\Delta h \ln \Theta + \sum_{i=1}^n \sigma_i^{ls} A_i^{ls}, \quad (2)$$

where $\Delta h < 0$ is the enthalpy of fusion (see, e.g., ref. 23), and $\Theta = T/T_0$. Note that in eq.(1) the mechanical effects within the crystal (e.g., stresses) are considered to reduce to an isotropic

pressure P^s , so that^{8,13}

$$P^s - P^l = \frac{2\sigma_i^{ls}}{h_i} \quad (i = 1, \dots, n), \quad (3)$$

where h_i is the distance from facet i to a point O within the crystal such that (see Figure 2)

$$\frac{\sigma_1^{ls}}{h_1} = \frac{\sigma_2^{ls}}{h_2} = \dots = \frac{\sigma_n^{ls}}{h_n}. \quad (4)$$

These equalities represent the necessary and sufficient conditions for the equilibrium shape of the crystal. This is known as the Wulff form and the equalities themselves are Wulff's relations (see, e.g., refs.8 and 13).

Equation (3) applied to the crystal is the equivalent of Laplace's equation applied to liquid. Thus, just as for a droplet, one can expect to find a high pressure within a small crystal. Using eqs.(3) and (4), one can show that, for a crystal surrounded by the liquid phase (melt),^{11,12}

$$V^s(P^s - P^l) = \frac{2}{3} \sum_{i=1}^n \sigma_i^{ls} A_i^{ls}. \quad (5)$$

By definition, the critical crystal (i.e., nucleus) is in equilibrium with the surrounding melt. For such a crystal the first term in eq.(1) vanishes. Therefore, by virtue of eq.(5), W_* , the reversible work of formation of the nucleus, can be written as

$$W_* = \frac{1}{2} V_*^s (P_*^s - P^l) = \frac{1}{3} \sum_{i=1}^n \sigma_i^{ls} A_{i*}^{ls}, \quad (6)$$

where the subscript “*” indicates quantities corresponding to the critical crystal (nucleus).

Now consider the case where a crystal cluster forms with one of its facets (say, facet λ) at a droplet surface. Assuming that $A_\lambda^{sv}/\pi R^2 \ll 1$, the deformation of the droplet can be neglected¹¹ and the reversible work of formation of a crystal particle with its facet λ at the droplet surface is

$$\widetilde{W} = \nu[\mu^s(P^s, T) - \mu^l(P^l, T)] - \widetilde{V}(P^s - P^l) + \sum_{i=1}^n \sigma_i^{ls} A_i^{ls} + \sigma_\lambda^{vs} A_\lambda^{vs} - \sigma^{vl} A_\lambda^{vs}, \quad (7)$$

where \tilde{V} is the volume of the crystal and where \sum^λ indicates that the term with $i = \lambda$ is excluded from the sum.

Wulff's relations in eq.(4), which determine the equilibrium shape of a crystal, can be regarded as a series of equilibrium conditions on the crystal "edges" formed by adjacent facets.

For example, on the edge between facets i and $i + 1$ the equilibrium condition is $\frac{\sigma_i^{ls}}{h_i} = \frac{\sigma_{i+1}^{ls}}{h_{i+1}}$.

In the case where one of the facets (facet λ) is the crystal-vapor interface while all the others lie within the liquid phase (see Figure 2), the equilibrium conditions on the edges formed by this facet with the adjacent ones (hereafter marked by a subscript j) are given by

$$\frac{\sigma_j^{ls}}{h_j} = \frac{\sigma_\lambda^{sv} - \sigma^{lv}}{\tilde{h}_\lambda}. \quad (8)$$

Note that the height of the λ -th pyramid (constructed with base on facet λ and with apex at point O of the Wulff crystal) will differ from that with all of the facets in the liquid. Thus, the shape of the crystal will differ from that in which all facets are in contact with the liquid. For this case, Wulff's relations take the form

$$\frac{\sigma_1^{ls}}{h_1} = \dots = \frac{\sigma_{\lambda-1}^{ls}}{h_{\lambda-1}} = \frac{\sigma_\lambda^{ls} - \sigma^{lv}}{\tilde{h}_\lambda} = \frac{\sigma_{\lambda-1}^{ls}}{h_{\lambda-1}} = \dots = \frac{\sigma_n^{ls}}{h_n}, \quad (9)$$

and eq.(3) becomes

$$P^s - P^l = \frac{2\sigma_i^{ls}}{h_i} \quad (i = 1, \dots, \lambda - 1, \lambda + 1, \dots, n), \quad P^s - P^l = \frac{2(\sigma_\lambda^{sv} - \sigma^{lv})}{\tilde{h}_\lambda}. \quad (10)$$

Making use of equations (9) and (10), one can represent eq.(7) as

$$\widetilde{W} = -\nu\Delta h \ln \Theta + \sum_{i=1}^n \sigma_i^{ls} A_i^{ls} + \sigma_\lambda^{vs} A_\lambda^{vs} - \sigma^{vl} A_\lambda^{vs}. \quad (11)$$

For a crystal with one of its facets a solid-vapor interface, and the others interfaced with the liquid, one can show that

$$\widetilde{V}(P^s - P^l) = \frac{2}{3} \left(\sum_{i=1}^n \sigma_i^{ls} A_i^{ls} + \sigma_\lambda^{vs} A_\lambda^{vs} - \sigma^{vl} A_\lambda^{vs} \right). \quad (12)$$

The reversible work \widetilde{W}_* of formation of a *critical* crystal can be thus represented as

$$\widetilde{W}_* = \frac{1}{2} \widetilde{V}_*(P_*^s - P^l) = \frac{1}{3} \left(\sum_{i=1}^{\lambda} \sigma_i^{ls} A_i^{ls} + \sigma_{\lambda}^{vs} A_{\lambda}^{vs} - \sigma_{\lambda}^{vl} A_{\lambda}^{vs} \right). \quad (13)$$

The similarity of equations (13) and (6) allows one to meaningfully compare them. One can show^{11,12} that the difference $P_*^s - P^l$ for the nucleus is determined exclusively by the degree of supercooling of the liquid, so that in both eq.(6) and eq.(13)

$$(P_*^s - P^l) = \frac{\Delta h}{v} \ln \Theta, \quad (14)$$

where v is the volume per molecule in a solid phase. Using eqs.(3) and (10), one obtains

$$\tilde{h}_{\lambda} = \frac{\sigma_{\lambda}^{vs} - \sigma_{\lambda}^{lv}}{\sigma_{\lambda}^{ls}} h_{\lambda}. \quad (15)$$

On the other hand, $h_i = \tilde{h}_i$ for $i = 1, \dots, \lambda - 1, \lambda + 1, \dots, n$, by virtue of eqs.(3), (10), and (14). Therefore, the Wulff shape of the surface-stimulated crystal is obtained by simply changing the height of the λ -th pyramid of the volume-based Wulff crystal. It is thus clear that if $\sigma_{\lambda}^{sv} - \sigma_{\lambda}^{lv} < \sigma_{\lambda}^{ls}$, then $\tilde{h}_{\lambda} < h_{\lambda}$ and hence $\widetilde{V}_* < V_*$. Since

$$\frac{\widetilde{W}_*}{W_*} = \frac{\widetilde{V}_*}{V_*} \quad (16)$$

(according to eqs.(6),(13) and (14)), we can conclude that if

$$\sigma_{\lambda}^{sv} - \sigma_{\lambda}^{lv} < \sigma_{\lambda}^{ls}, \quad (17)$$

then $\widetilde{W}_* < W_*$. In other words, if condition (17) is fulfilled, it is thermodynamically more favorable for the crystal nucleus to form with facet λ *at* the droplet surface rather than *within* the droplet.

Inequality (17) coincides with the condition of partial wettability of the λ -th facet of the crystal by its own liquid phase¹³ (note that it has exactly the same form in multicomponent

droplets¹²). This result is physically reasonable, because,²⁴ if the condition of partial wettability holds, the free energy per unit area required to form a *direct* interface between bulk vapor and solid (as in case of surface-stimulated crystallization) is less than the free energy required to form a uniform *intruding* layer of liquid phase, which involves creation of two interfaces “solid-liquid and “liquid-vapor”. This intuitive argument alone, however, is *not sufficient* to claim that the droplet surface stimulates crystal nucleation, because (as shown in ref.11 and 12) the volume-located and surface-faced nuclei *differ* from each other in shape and size (and, possibly, composition). It is necessary to compare the total surface contributions to the free energy of nucleus formation rather than the free energies per unit area of one particular facet of the nucleus, i.e., it is necessary to compare the quantity $\frac{1}{3} \left(\sum_{i=1}^n {}^\lambda \sigma_i^{ls} A_i^{ls} + \sigma_\lambda^{sv} A_\lambda^{sv} - \sigma^{lv} A_\lambda^{sv} \right)$ to $\frac{1}{3} \sum_{i=1}^n \sigma_i^{ls} A_i^{ls}$ rather than quantity $\sigma_\lambda^{sv} A_\lambda^{sv} - \sigma_\lambda^{lv} A_\lambda^{sv}$ to $\sigma_\lambda^{ls} A_\lambda^{ls}$. It is a mere coincidence that the first comparisons reduces to the second one.

3 Kinetics of surface-stimulated crystal nucleation

Inequality (17) allows one to predict whether crystallization in a supercooled droplet will or will not be thermodynamically stimulated by the surface. To apply this in practice, however, one needs accurate and detailed information about the surface tension of the liquid-vapor interface and the surface tensions of crystal facets both in the liquid and in the vapor. Data on σ^{lv} are available for most liquids of interest or can be easily obtained. The availability of data on σ^{ls} and σ^{sv} is more problematic. Data on σ^{ls} are often obtained by matching experimental crystal nucleation rates with the predictions of CNT, treating the surface tension of the crystal nucleus as an adjustable parameter.^{9,25} However, such data are not suitable for using in eq.(17) for several reasons (see ref.11,12 for more details), one of which is the unsatisfactory state of the kinetic theory of crystal nucleation in light of the results outlined in the previous section.

Indeed, the classical expression for the rate of crystal nucleation conventionally used in atmospheric models as well as for treating experimental data, is derived by assuming that crystal nuclei form *within* the liquid.^{8,9,2,3} However, under conditions of partial wettability of at least one crystal facet by its melt, the formation of a crystal nucleus with that facet at the droplet surface is thermodynamically favored over its formation with all the facets *within* the droplet. This effect can become important when the crystallizing liquid is in a dispersed state, which is the case with the freezing of atmospheric droplets^{2,3} and many experiments.²⁶

Assuming a monodisperse (or sharp enough Gaussian) distribution of liquid droplets, the average crystallization time of the ensemble equals that of a single droplet (for simplicity, hereafter we will discuss only the monodisperse distribution, although results will be also applicable to a narrow enough Gaussian-like distribution). Let us denote that time by t_1 . Typical sizes of atmospherically relevant droplets allow one to assume that the formation of a single crystal nucleus in a droplet immediately leads to the crystallization of the latter, i.e., the time of growth of a crystal nucleus to the size of the whole droplet is negligible compared to the time necessary for the first nucleation event in the droplet to occur (in experiments this can be achieved by using special techniques²⁶). Consequently,

$$t_1 = 1/I, \quad (18)$$

where I is the per-particle (pp) nucleation rate, i.e., the total number of crystal nuclei appearing in the whole volume of the liquid droplet per unit time. Until recently, atmospheric models had considered homogeneous crystal nucleation in droplets to be exclusively volume-based, with

$$I = I^{vb} \equiv J_v V_1, \quad (19)$$

where V_1 is the volume of a single droplet and J_v is the the rate of volume-base crystal

nucleation given, e.g., by^{9,2,3}

$$J_v = \frac{kT}{h} \rho_l e^{-W_*/kT}, \quad (20)$$

with k and h being the Boltzmann and Planck constants and ρ_l the number density of molecules in the liquid phase. However, since the surface-to-volume ratio of a droplet increases with decreasing droplet size, eq.(19) may become inadequate for small enough droplets in which the surface-stimulated crystal nucleation can compete with (or even dominate) the volume-based process. To take this possibility into account, it was recently^{19,20} suggested that instead of eq.(19) the pp-rate of crystal nucleation should consist of two contributions,

$$I = J_s S_1 + J_v V_1, \quad (21)$$

where S_1 is the area of the droplet surface and J_s is the number of crystal nuclei forming per unit time on unit surface area of the droplet (i.e., in a surface stimulated mode). The rate J_s was conjectured^{19,20} to have the following form (reminiscent of that for the rate of crystal nucleation on heterogeneous surfaces^{2,3})

$$J_s = \frac{kT}{h} \rho_l^s e^{-\tilde{W}_*/kT}, \quad (22)$$

with ρ_l^s being the number of (liquid phase) molecules per unit area of the droplet surface.

As clear, using eq.(22) for J_s on the RHS of eq.(21) amounts to implicitly assuming that in the surface-stimulated mode the crystal nucleus forms in a heterogeneous fashion on a molecule m_s located in the droplet surface monolayer. As a first step, the first nearest neighbors (both in the bulk and in the droplet surface layer) of molecule m_s would acquire a stable crystalline configuration (this would be due to fluctuations and hence might be temporary). At the second step, the second nearest neighbors (including those in the droplet surface layer) of molecule m_s would acquire a stable crystalline order (this would be again due to fluctuations and hence temporary). These steps would continue until the crystalline cluster formed around

m_s attained a critical size (i.e., became a crystal nucleus with one of its facets at the droplet-vapor interface) which would be followed by a quick crystallization of the whole droplet. The number of nuclei forming per unit time per unit area of the droplet surface (i.e., the surface nucleation rate J_s) would be then given by eq.(22).

However, there is a weak point in the above reasoning. Indeed, heterogeneous mechanism of nucleation implies that the *initial* stage of the formation of a new phase fragment around a heterogeneous center is thermodynamically favorable, i.e., it is accompanied by a decrease in the appropriate free energy of the system. This was clearly demonstrated for heterogeneous condensation on ions^{27,28} as well as on insoluble,^{29,30} mixed,³¹ and various soluble³² macroscopic particles. For all these phenomena a specific physical effect, causing the initial decrease in the free energy of the system upon the formation of a new phase particle around a heterogeneous center, can be unambiguously identified. This is not the case with surface-stimulated crystal nucleation in droplets. The formation of a crystal nucleus with one of its facets at the droplet surface cannot *start* preferentially *at* the surface, because the latter does not have any sites which would make the ordering of the surrounding *surface located* molecules thermodynamically more favorable than the ordering of interior molecules. On the contrary, the surface layer of a crystalline structure remains disordered far below the freezing/melting temperature. This phenomenon is often referred to as premelting³³ and was observed both experimentally^{34–36,15} and by simulations^{37–40}.

According to the empirical criterion proposed by Lindemann,⁴¹ melting in the bulk can occur when the root mean amplitude of thermal vibrations of an atom exceeds a certain threshold value of more or about 10% of the distance to the nearest neighbor in the crystalline structure. Developing this criterion, Tammann suggested³³ that the outermost layer of the crystal should become disordered far below the bulk melting point due to the higher freedom

of motion for surface-located molecules which have a reduced number of neighbors and hence have a higher vibrational amplitude compared to those in the bulk. Thus, one can expect the Lindemann criterion for surface atoms to be satisfied at a temperature lower than that for atoms located within the crystalline structure. The surface melting (often referred to as premelting) involves the formation of a thin disordered layer at a temperature significantly below the melting one.

Experimentally the premelting phenomenon was apparently first detected by Lyon and Somorjai³⁴ who studied the structures of clean (111), (110), and (100) crystal faces of platinum as a function of temperature by means of low-energy electron diffraction and observed the formation of disordered surface structures at temperatures far below the melting temperature $T_m = 2043$ K. Direct experiments on the surface-initiated melting were also carried out by Frenken *et al.*³⁵ using Rutherford backscattering in conjunction with ion-shadowing and blocking. That experiment revealed a reversible order-disorder transition on the (110) surface of a lead crystal well below its melting point $T_m = 600.7$ K. Since then, other techniques have been employed such as calorimetry, electron, neutron, and X-ray diffraction, microscopy, ellipsometry, and helium scattering. Although most experiments were carried out under equilibrium conditions, melting tended to be initiated at the surface even when the crystalline solid was heated very quickly so that equilibrium conditions were not established.³⁶ Lately, molecular dynamics simulations have been also widely used to study premelting (see, e.g., ref.37 and recent simulations of premelting in *AgBr* (ref.38), in *Cr₂O₃* (ref.39), and the premelting of a clean *Al*(110) surface⁴⁰).

Of utmost atmospheric relevance, there has been accumulated undeniable evidence^{16,42–44} for the premelting of ice (first apparently discussed by Faraday⁴⁵). Relatively recently Wei *et al.*⁴⁶ experimentally observed that the premelting of the (0001) face of hexagonal ice occurs

at the temperature of about 200 K, i.e., much below the lowest temperature reported for homogeneous freezing of atmospheric droplets. Thus, although the droplet surface can (under condition (17)) stimulate crystal nucleation, a crystal nucleus with one facet as a droplet-vapor interface most likely *begins* its formation (as a subcritical crystal) homogeneously in a spherical layer adjacent to the droplet surface (“sub-surface layer”). When this crystal becomes large enough (due to fluctuational growth usual for the nucleation stage), one of its facets hits the droplet surface and at this moment or shortly thereafter it becomes a nucleus owing to a drastic change in its thermodynamic state.

Let us consider the “sub-surface” layer of the droplet hereafter referred to as an SSN layer (SSN can stand for both “sub-surface nucleation” and “surface-stimulated nucleation”). Its thickness will be denoted by η (more detailed discussion of η is given in the following subsection). By definition (albeit somewhat loose for now), any crystalline cluster that starts its evolution with its center in the SSN layer has a potential to become a nucleus (by means of structural and density fluctuations) once one of its facets that satisfies the condition of partial wettability, eq.(17), meets the droplet surface. Clearly, to become a surface-stimulated nucleus, the subcritical cluster must evolve in such a way that its facet λ (satisfying condition (17)) is parallel to the droplet surface at the time they meet. The orientation adjustment cannot be mechanical because this would require relatively long time scales, but may, or may not, occur by means of appropriate spatial distribution of density and structure fluctuations around the cluster.

In the framework of the SSN layer model, the pp-rate of crystal nucleation is given by the sum

$$I = J_v^s V_1^s + J_v(V_1 - V_1^s) \quad (23)$$

(rather than by eqs.(19) or (21)), where J_v^s is the number of crystal nuclei forming in a surface

stimulated mode per unit time in unit volume of the SSN layer whereof the total volume is V_1^s . Equivalently,

$$I = I^{ss} + I^{vb}, \quad (24)$$

where

$$I^{ss} = (J_v^s - J_v)V_1^s \quad (25)$$

and I^{vb} (defined by eq.(19)) are the contributions to the total pp-rate of crystal nucleation arising from the surface-stimulated and volume-based modes, respectively. As clear from the above discussion,

$$J_v^s = \frac{kT}{h} \rho_l e^{-\tilde{W}_*/kT}, \quad (26)$$

where the number density of molecules in the droplet is assumed to be uniform up to the dividing surface,¹³ in consistency with the capillarity approximation.¹⁰ In this approximation, the pre-exponential factors in eqs.(20) and (26) are the same, which reflects the *homogeneous* nature of the nascent of a sub-critical cluster in both volume-based and surface-stimulated modes of crystal nucleation. On the other hand, the nuclei in these two modes are different hence have different free energies of formation, which results in different exponents: $-W_*/kT$ for the volume-based process and $-\tilde{W}_*/kT$ for the surface-stimulated mode (see section 2).

By virtue of eqs.(19),(20), and (26), one can rewrite equation (25) as

$$I^{ss} = \alpha I^{vb}, \quad (27)$$

where

$$\alpha \equiv \alpha(\varepsilon) \equiv \alpha(\varepsilon, \Delta W_*(\eta(ep))) = [1 - (1 - \varepsilon)^3](-1 + e^{-\Delta W_*/kT}) \quad (28)$$

with $\varepsilon = \eta/R$ ($0 \leq \varepsilon \leq 1$). As clear from eq.(27), the ratio $\alpha = I^{ss}/I^{vb}$ characterizes the relative intensity of surface-stimulated and volume based modes of crystal nucleation in the droplet. If the volume-based process predominates over the surface-stimulated nucleation

(i.e., $I^{ss} \ll I^{vb}$), then $\alpha \ll 1$. If the surface-stimulated mode prevails over the volume-based one (i.e., $I^{ss} \gg I^{vb}$), then $\alpha \gg 1$. In the cross-over regime, when the nucleation mode factor is roughly in the range $0.33 \lesssim \alpha \lesssim 3$, the contributions from both modes to the total pp-rate of crystal nucleation are comparable to each other. The exact cross-over point is given by the equality $I^{ss} = I^{vb}$.

According to eq.(28), under given external conditions (temperature, pressure, etc...) the value of the “nucleation mode factor” α is determined by $\varepsilon = \eta/R$, i.e., by the size of the freezing droplet, R . The geometric factor $1 - (1 - \varepsilon)^3$ in α monotonically increases from 0 to 1 with increasing ε . However, the nucleation mode factor itself, α , may have a more complicated dependence on ε because of the exponential $e^{-\Delta W_*/kT}$ in which ΔW_* is intrinsically related to ε via η .

3.1 Nucleation mode factor

According to the definition of the SSN layer, its thickness, denoted by η , is determined by the shape and orientation of the crystal nucleus and the physical characteristics of the crystal nucleus *and* droplet. They also determine the free energy of nucleus formation in both surface-stimulated and volume-based modes, \widetilde{W}_* and W_* , respectively (the droplet surface tension is involved only in \widetilde{W}_* but not in W_*). For a given droplet in a given thermodynamic state, η and W_* are completely independent of each other. However, both η and \widetilde{W}_* are determined by the shape *and* orientation of the crystal nucleus. Hence the dependence $\Delta W_* = \Delta W_*(\eta)$ which, according to eq.(28), is likely to be a key factor in determining the ε -dependence of I^{ss} (because $\varepsilon = \eta/R$) and, ultimately, the value of the nucleation mode factor α .

Consider first the case where crystal clusters (including the critical one, nucleus, which forms as a result of fluctuational growth of an initially subcritical cluster) have n facets each and

assume that only one of these facets, say, facet λ , satisfies the condition of partial wettability, eq.(17). Let us introduce the unit vectors \mathbf{n}_d and \mathbf{n}_λ as the external normal vectors to the droplet surface and facet λ , respectively. The angle Θ between \mathbf{n}_d and \mathbf{n}_λ determines the mutual orientation of the droplet surface and facet λ . Clearly, $0 \leq \Theta \leq \pi$ with $\Theta = 0$ corresponding to the λ -facet being parallel to the droplet surface and $\Theta = \pi$ their being “antiparallel”.

If the only possible orientation of a crystal cluster were the one with $\Theta = 0$, then the λ -facet of any surface-stimulated nucleus would be a part of the droplet surface. Consequently, the thickness η of the SSN layer would be equal to \tilde{h}_λ (the height of the λ^{th} pyramid, with facet λ its basis and point O its apex, see Figure 2) and eq.(27) could be written as

$$I^{ss} = [1 - (1 - \varepsilon_\lambda)^3](-1 + e^{-\Delta W_{*\lambda}/kT})I^{vb}, \quad (29)$$

where $\varepsilon_\lambda = \tilde{h}_\lambda/R$, $\Delta W_{*\lambda} \equiv \tilde{W}_{*\lambda} - W_*$, and $\tilde{W}_{*\lambda}$ is the free energy of formation of a surface-stimulated nucleus with facet λ being a part of the droplet surface.

In reality, however, the orientations of crystal clusters are randomly distributed in the range from $0 \leq \Theta \leq \pi$. Assuming this distribution to be uniform (because there are no obvious reasons for the contrary), its probability density is $p(\Theta) = \frac{1}{\pi}$ with the normalization $\int_0^\pi d\Theta p(\Theta) = 1$.

Let $p_{\Theta\varepsilon}(\Theta)$ be the probability density that a crystal cluster has an orientation Θ and, at this orientation, the surface-stimulated nucleation occurs in a layer of (dimensionless) thickness ε . As mentioned above, for a given droplet under given external conditions ε can be a function of only Θ . Moreover, the droplet surface can stimulate the nucleation of crystal clusters only at one single orientation, $\Theta = 0$. Thus,

$$p_{\Theta\varepsilon}(\Theta) = p(\Theta)\delta(\Theta) = \frac{1}{\pi}\delta(\Theta). \quad (30)$$

The contribution I^{ss} from the surface-stimulated mode to the total pp rate of crystal nucleation, I , is now obtained by averaging $I^{ss}(\varepsilon(\Theta))$, given by eqs.(27) and (28), over all the possible orientations of crystal clusters, i.e., as

$$I^{ss} = \int_0^\pi d\Theta [1 - (1 - \varepsilon(\Theta))^3](-1 + e^{-\Delta W_*(\Theta)/kT}) I^{vb} p_{\Theta\varepsilon}(\Theta). \quad (31)$$

Taking into account eq.(30) and the equality $\varepsilon(0) = \varepsilon_\lambda$, one obtains

$$I^{ss} = \frac{1}{\pi} [1 - (1 - \varepsilon_\lambda)^3](-1 + e^{-\Delta W_{*\lambda}/kT}) I^{vb}, \quad (32)$$

As clear, the ability of a crystal cluster in the SSN layer to appear with facet λ not only parallel to the droplet surface but with any other orientation, with Θ uniformly distributed from 0 to π , decreases I^{ss} (and hence α) by a factor of $1/\pi$ compared to a hypothetical situation when all crystal clusters would evolve with their λ facets parallel to the droplet surface, i.e., with $\Theta = 0$.

Equation (32) is obtained for the case where every cluster of the nascent crystalline structure has only one facet (facet λ) satisfying the condition of partial wettability, eq.(17). In a more general situations, every cluster can have w facets ($1 \leq w \leq N$) partially wettable by the melt. Let these facets be numbered 1 through w . Every one of these facets contributes to the surface-stimulated mode of the pp-rate of crystal nucleation in the droplet. Since these contributions I_λ^{ss} ($\lambda = 1, \dots, w$) are independent of one another, each of them is determined by eq.(32), so that the total “surface-stimulated” contribution I^{ss} to the pp-rate of crystal nucleation will be given by the sum $\sum_{\lambda=1}^w I_\lambda^{ss}$, i.e.,

$$I^{ss} = \sum_{\lambda=1}^w \frac{1}{\pi} [1 - (1 - \varepsilon_\lambda)^3](-1 + e^{-\Delta W_{*\lambda}/kT}) I^{vb\lambda}. \quad (33)$$

In a rough approximation, one can assume that

$$\tilde{h}_0 \equiv \tilde{h}_1 \approx \tilde{h}_2 \approx \dots \approx \tilde{h}_w \quad (34)$$

and

$$\widetilde{W}_{*0} \equiv \widetilde{W}_{*1} \approx \widetilde{W}_{*2} \approx \dots \approx \widetilde{W}_{*w}. \quad (35)$$

The former assumption is reasonable if, for instance, all crystal nuclei have globular (not elongated) shape with aspect ratios close to 1, whereas the latter implies that the surface tensions of facets $1, \dots, w$ do not differ much from one another. With such approximations, equation (33) reduces to

$$I^{ss} = \frac{w}{\pi} [1 - (1 - \varepsilon_0)^3] (-1 + e^{-\Delta W_{*0}/kT}) I^{vb}, \quad (36)$$

where $\varepsilon_0 = \tilde{h}_0/R$ and $\Delta W_{*0} = \widetilde{W}_{*0} - W_*$. This equation is convenient for rough numerical evaluations of I^{ss} . In a more complicated case where assumption (34) is not acceptable, while approximate equalities in eq.(35) do hold, eq.(33) acquires the form

$$I^{ss} = \frac{1}{\pi} [w - \sum_{\lambda=1}^w (1 - \varepsilon_\lambda)^3] (-1 + e^{-\Delta W_{*0}/kT}) I^{vb} \dots \quad (37)$$

where the factor in the square brackets is a relatively weak function of ε_λ ($\lambda = 1, \dots, w$) not exceeding w . Again, this form is more convenient than eq.(33) to numerically evaluate the nucleation mode factor α .

4 Numerical Evaluations and Experimental Perspective

To illustrate the above theory by numerical evaluations, consider the freezing of water droplets (surrounded by water vapor in air) at around $T = 233$ K (i.e., about -40°C). The homogeneous and isothermal character of freezing is assumed.

As reported by Defay *et al.*¹³, the rate of crystal nucleation in bulk supercooled water at this temperature is $7 \times 10^{12} \text{ cm}^{-3} \text{ s}^{-1}$, with the nucleation barrier height $W_* = 45 \text{ kT}$, the average (over all crystal facets) surface tension of liquid-solid (water-ice) interface σ^{ls} being

about 20 dyn/cm (Table 18.1 in ref.12). The surface tensions of liquid-vapor and solid-vapor (ice-water vapor) interfaces at $T = 233$ K will be taken to be $\sigma^{lv} = 88$ dyn/cm and $\sigma^{sv} = 103$ dyn/cm, respectively. All these values of σ^{ls} , σ^{lv} , and σ^{sv} are consistent with the data provided in ref.2.

The wettability of a solid by a liquid (both in contact with a vapor) is determined by the contact angle, defined as the angle between the tangents to the liquid-vapor and solid-liquid interfaces at the three phase contact line. According to Young's relation,¹³ which gives a connection between three interfacial tensions and contact angle, the above values of σ^{ls} , σ^{lv} , and σ^{sv} would correspond to the contact angle $\beta \simeq 19.4^\circ$ (or $\cos \beta \simeq 0.943$). Therefore, at $T = 233$ K at least some of (if not all) the facets of an ice crystal are only partially wettable by liquid water. This is consistent with the experimental results of Elbaum *et al.*¹⁶ who reported partial wettability of the basal facets of hexagonal ice (Ih) at temperatures slightly below 0°C. In those experiments,¹⁶ when air was added to water vapor the partial wetting of ice by water transformed into complete wetting but only for some orientations. Besides, the wettability of solids by fluids usually decreases with decreasing temperature,^{47,48} so one can expect that at temperatures far below 0°C at least some facets of water crystals remain only partially wettable even in the presence of air. Moreover, according to Cahn's theory,⁴⁷ perfect wetting of a solid by a liquid away from the critical point is not generally observed, i.e., condition (17) should be fulfilled for any substance at sufficiently low temperatures. That theory⁴⁷ can be also applied to the case where the solid is of the same chemical nature as the fluid phases (it is only assumed that the surface of the solid phase is sharp on an atomic scale and interactions between surface and fluid are sufficiently short-range). Cahn's theory is inapplicable at temperatures close to the fluid critical point, but temperatures involved in crystallization are usually far below that point.

For simplicity of numerical evaluations, let us assume that only the basal facet $\{0001\}$ of the hexagonal ice crystal is partially wettable by water at $T = 233$ K. Denote the height of the basal pyramid of the crystal cluster by \tilde{h}_b when the basal facet is at the droplet surface and by h_b when the entire crystal is immersed in the droplet. According eq.(15),

$$\frac{\tilde{h}_b}{h_b} = \frac{95 - 80}{20} = 3/4. \quad (38)$$

Consequently, the ratio of the volumes of the surface-based and volume-formed clusters is

$$\frac{\widetilde{V}_*}{V_*} \simeq \frac{\tilde{h}_b + h_b}{2h_b}. \quad (39)$$

By virtue of eq.(16), for the corresponding works of formation of crystal nuclei we obtain

$$\frac{\widetilde{W}_*}{W_*} \simeq 0.875, \quad (40)$$

i.e., $\widetilde{W}_* \simeq 39.3 kT$. Thus, the decrease in the work of formation of the surface-stimulated crystal nucleus, as compared to that of the volume-based one, is $\Delta W_* \approx -5.7 kT$.

Let us now evaluate the thickness of the SSN layer, η , which still depends on the size and habit of crystal nuclei even after averaging over Θ in eq.(31). According to eq.(6), the work of formation of a volume-formed crystal nucleus can be rewritten in the form

$$W_* = \frac{1}{3}\sigma_{ls}A_*, \quad (41)$$

where σ_{ls} is the average (over all the crystal facets) interfacial tension of the nucleus and A_* is its total surface area.^{11,12} The total surface area of an Ih crystal (shaped as a right prism) is $A = 12h_b a + 3\sqrt{3}a^2$, where a is the side length of the basal facet (for a regular hexagon a is also equal to its radius) and h_b is the half-height of the prism.

One can define the aspect ratio of a volume-formed Ih crystal as $\gamma = h_b/a$. For the formation of snow crystals from water vapor γ is a complex non-monotonic function of both temperature and water vapor saturation ratio.⁴⁹ Likewise, for crystallization in liquid water

γ is a complex function of temperature and pressure.⁵⁰ For this reason, let us consider two opposite cases, $\gamma = 2$ and $\gamma = 0.5$, corresponding to column-like and plate-like crystals of Ih , respectively. As mentioned, for crystallization in pure water $W_* \approx 45 kT$ and $\sigma_{ls} \approx 20 \text{ dyn/cm}$ at $T = 233 \text{ K}$. Therefore, eq.(41) leads to $h_b \approx 6.9 \times 10^{-8} \text{ cm}$ for $\gamma = 0.5$ and $h_b \approx 17.3 \times 10^{-8} \text{ cm}$ for $\gamma = 2$ which, according to eq.(38), correspond to the following values of the thickness of the SSN layer:

$$\eta = \tilde{h}_b \approx 5.2 \times 10^{-8} \text{ cm} \quad (\gamma = 0.5), \quad \eta = \tilde{h}_b \approx 12.9 \times 10^{-8} \text{ cm} \quad (\gamma = 2). \quad (42)$$

The nucleation mode factor α can be estimated from eq.(36) with $w = 2$. Its dependence on the radius of the droplet is presented in Figure 3, where the dashed and solid curves correspond to $\gamma = 2$ and $\gamma = 0.5$, respectively. As clear, the surface-stimulated mode can considerably enhance crystal nucleation in droplets with radii even exceeding $1 \mu\text{m}$. For example, for droplets of radius $R = 0.2 \mu\text{m}$ the nucleation mode factor $\alpha = 3.3$ for $\gamma = 2$ and $\alpha = 1.4$ for $\gamma = 0.5$, while for droplets of radius $R = 2 \mu\text{m}$ the nucleation mode factor $\alpha = 0.34$ for $\gamma = 2$ and $\alpha = 0.14$ for $\gamma = 0.5$. These estimates suggest that homogeneous crystal nucleation in water droplets with radii smaller than $R \approx 0.2 \mu\text{m}$ occurs predominantly in the surface-stimulated mode, while the volume-based mode prevails in droplets with radii greater than $R \approx 2 \mu\text{m}$. Both the surface-stimulated and volume-based modes apparently provide contributions of the same order of magnitude to the total pp-rate of crystal nucleation in droplets with radii in the range approximately from $0.2 \mu\text{m}$ to $2 \mu\text{m}$. Similar evaluations can be carried out for the case where all the facets of Ih crystals are only partially wettable. Assuming that for the Ih prism facets ΔW_* is approximately the same as for the basal ones, it is clear from eq.(36) that in this case the nucleation mode factor α may be greater than 1 even for droplets with $R > 10 \mu\text{m}$ (i.e., the surface-stimulated mode may dominate even for such relatively large droplets).

While the above numerical estimates are approximate because of insufficiently accurate data on the interfacial tensions involved in the model, laboratory techniques currently available for studying crystal nucleation in droplets make it possible to carry out the experimental verification of the above theory. Indeed, modern experimental methods^{2,26} can provide data on the dependence of the pp-rate of crystal nucleation in droplets on their radius, i.e., I as function of R . According to the above model, one can expect that there must exist such a constant A that for $R \gtrsim 1 \mu\text{m}$ the LMS fit of the experimental dependence I^{\exp} vs R with the function $(1 + A/R)BR^3$ (B is another constant) is much more accurate than with the function BR^3 and the inaccuracy of the latter compared to the former should be aggravating with decreasing R .

5 Concluding Remarks

The thermodynamics of surface-stimulated crystal nucleation was previously developed for both unary¹¹ and multicomponent droplets¹² (for which the theory is more complicated not only due to the presence of several components, but also due to the surface adsorption of all components as well as their dissociation into ions) in the framework of CNT. A criterion was found for when the surface of a droplet can stimulate crystal nucleation therein so that the formation of a crystal nucleus with one of its facets at the droplet surface is thermodynamically favored (i.e., occurs in a surface stimulated mode) over its formation with all the facets *within* the liquid phase (i.e., in a volume-based mode). For both unary and multicomponent droplets, this criterion coincides with the condition of partial wettability of at least one of the crystal facets by the melt. However, so far the kinetic aspects of this phenomenon had not been studied at all.

In this paper we have presented a kinetic theory of homogeneous crystal nucleation in unary

droplets taking into account that a crystal nucleus can form not only in the volume-based mode but also in the surface-stimulated one. We have invoked experimental and simulations-based evidence showing that surface-stimulated crystal nucleation is *not* a particular case of heterogeneous nucleation. On the contrary, it is a particular case of homogeneous crystal nucleation hence its thermodynamic similarities with heterogeneous nucleation can be misleading because the kinetics of this process cannot be treated by using the formalism of heterogeneous nucleation on foreign surfaces.

Even in the surface-stimulated mode the crystal nucleus initially emerges (as a subcritical cluster) homogeneously in the droplet sub-surface layer, not pseudo-heterogeneously at the droplet surface. A homogeneously emerged sub-critical crystal cluster can become a surface-stimulated nucleus when, after growing large enough owing to density and structure fluctuations, one of its facets meets the droplet surface and both are parallel to each other. This effect gives rise to an additional contribution to the total rate of crystal nucleation in a droplet (the conventional contribution arises from the volume-based crystal nucleation). We have derived an expression for the total per-particle rate of crystal nucleation in the droplet in the framework of CNT. The theory has been presented only for unary droplets, but its generalization to multicomponent droplets is possible although not straightforward.

As a numerical illustration of the proposed theory, we have considered crystal nucleation in water droplets at $T = 233$ K. Our results suggest that that the surface-stimulated mode can markedly enhance the per-particle rate of crystal nucleation in water droplets as large as $10 \mu\text{m}$ in radius. We have also roughly outlined a simple way to carry out the experimental verification of the proposed theory.

However complex a theory of homogeneous crystal nucleation in droplets may be, the presence of foreign particles, serving as nucleating centers, makes the crystal nucleation phe-

nomenon (and hence its theory) even more involved. Numerous aspects of heterogeneous crystal nucleation still remain obscure. For example, it has been observed that the same nucleating center initiates the crystallization of a supercooled droplet at a higher temperature in the contact mode (with the foreign particle just *touching* the droplet surface) than in the immersion mode (particle *immersed* in the droplet).[2,5,21,22] Underlying physical reasons for this enhancement have remained unclear, but as little as might be known about the phenomenon of surface-stimulated (homogeneous) crystal nucleation, it strongly suggests that the droplet surface can enhance heterogeneous nucleation in a way similar to the enhancement of the homogeneous process. The thermodynamics and kinetics of heterogeneous crystal nucleation in droplets (in both contact and immersion modes) is the subject of our current research.

Acknowledgment - YSD thanks Raymond Shaw for helpful discussions.

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Captions

to Figures 1 to 3 of the manuscript

EFFECT OF THE SURFACE-STIMULATED MODE ON THE KINETICS OF HOMOGENEOUS CRYSTAL

NUCLEATION IN DROPLETS

by **Y. S. Djikaev**

Figure 1. A liquid droplet surrounded by vapor. Case *a*: homogeneous crystal nucleation *within* the droplet (volume-based mode). Case *b*: homogeneous crystal nucleation with one of crystal facets *at* the droplet surface (surfacc-stimulated mode).

Figure 2. Illustration to Wulff's relations (4) and (9). The surface area and surface tension of the facet *i* are denoted by A_i and σ_i , respectively; h_i is the distance from the facet *i* to the reference point O (see the text for more detail).

Figure 3. The nucleation mode factor $\alpha = I^{ss}/I^{vb}$ (given by eq.(36) with $w = 2$) as a function of R for crystal nucleation in water droplets at $T = 233$ K. The dashed and solid curves correspond to $\gamma = 2$ and $\gamma = 0.5$, respectively.

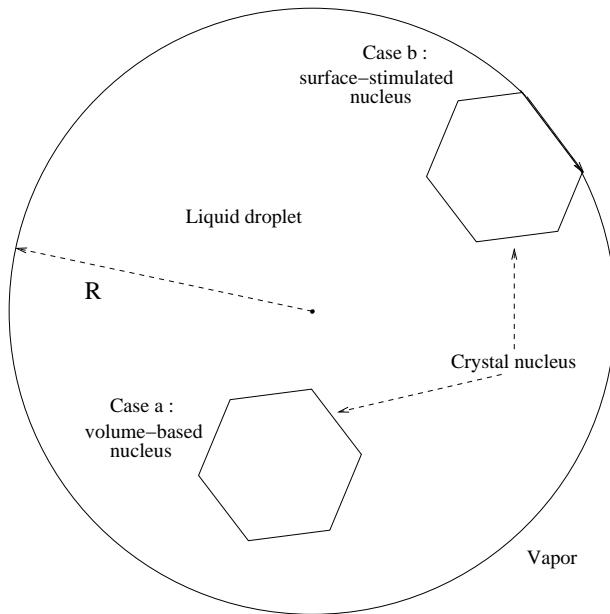


Figure 1:

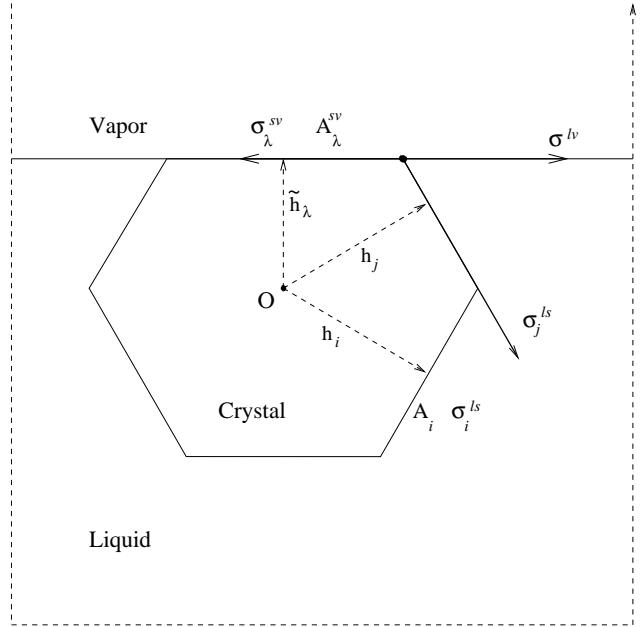


Figure 2:

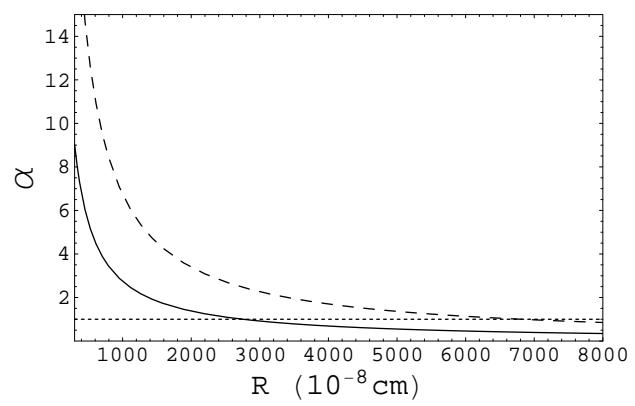


Figure 3: